I. Multiple Choice Questions (Type-I)

1. On addition of conc. H_2SO_4 to a chloride salt, colourless fumes are evolved but in case of an iodide salt, violet fumes come out. This is because
(i) H_2SO_4 reduces HI to I_2
(ii) HI is of violet colour
(iii) HI gets oxidised to I_2
(iv) HI changes to HIO_3
Solution:
Option (iii) is the answer.

2. In qualitative analysis when H_2S is passed through an aqueous solution of salt acidified with dil. HCl, a black precipitate is obtained. On boiling the precipitate with dil. HNO_3, it forms a solution of blue colour. Addition of excess of aqueous solution of ammonia to this solution gives _____.
(i) a deep blue precipitate of $Cu(OH)_2$
(ii) a deep blue solution of $[Cu(NH_3)_4]^{2+}$
(iii) a deep blue solution of $Cu(NO_3)_2$
(iv) a deep blue solution of $Cu(OH)_2.Cu(NO_3)_2$
Solution:
Option (ii) is the answer.

3. In a cyclotrimetaphosphoric acid molecule, how many single and double bonds are present?
(i) 3 double bonds; 9 single bonds
(ii) 6 double bonds; 6 single bonds
(iii) 3 double bonds; 12 single bonds
(iv) Zero double bonds; 12 single bonds
Solution;
Option (i) is the answer.

4. Which of the following elements can be involved in $p\pi-d\pi$ bonding?
(i) Carbon
(ii) Nitrogen
(iii) Phosphorus
(iv) Boron
Solution:
Option (iii) is the answer.

5. Which of the following pairs of ions are isoelectronic and isostructural?
(i) CO_3^{2-}, NO^{3-}
(ii) ClO^{3-}, CO_3^{2-}
(iii) SO_3^{2-}, NO^{3-}
(iv) ClO^{3-}, SO_3^{2-}
Solution:
Option (i) is the answer.

6. Affinity for hydrogen decreases in the group from fluorine to iodine. Which of the halogen acids should have the highest bond dissociation enthalpy?
(i) HF
(ii) HCl
(iii) HBr
(iv) HI
Solution:
Option (i) is the answer.

7. Bond dissociation enthalpy of E—H (E = element) bonds are given below. Which of the compounds will act as the strongest reducing agent?

Compound	**NH3**	**PH3**	**AsH3**	**SbH3**
Δdiss (E—H)/kJ mol–1	**389**	**322**	**297**	**255**

(i) NH3
(ii) PH3
(iii) AsH3
(iv) SbH3
Solution:
Option (iv) is the answer.

8. On heating with concentrated NaOH solution in an inert atmosphere of CO_2, white phosphorus gives a gas. Which of the following statement is incorrect about the gas?
(i) It is highly poisonous and has smelled like rotten fish.
(ii) It's a solution in water decomposes in the presence of light.
(iii) It is more basic than NH3
(iv) It is less basic than NH3
Solution:
Option (iii) is the answer.

9. Which of the following acids forms three series of salts?
(i) H_3PO_2
(ii) H_3BO_3
(iii) H_3PO_4
(iv) H_3PO_3
Solution:
Option (iii) is the answer.

10. Strong reducing behaviour of H_3PO_2 is due to
(i) The low oxidation state of phosphorus
(ii) Presence of two –OH groups and one P–H bond
(iii) Presence of one –OH group and two P–H bonds
(iv) High electron gain enthalpy of phosphorus
Solution:
Option (iii) is the answer.

11. On heating lead, nitrate forms oxides of nitrogen and lead. The oxides formed are _____.
(i) N_2O, PbO
(ii) NO_2, PbO
(iii) NO, PbO
(iv) NO, PbO_2
Solution;
Option (ii) is the answer.

12. Which of the following elements does not show allotropy?
(i) Nitrogen
(ii) Bismuth
(iii) Antimony
(iv) Arsenic
Solution:
Option (i) is the answer.

13. Maximum covalency of nitrogen is _____.
(i) 3
(ii) 5
(iii) 4
(iv) 6
Solution:
Option (iii) is the answer.

14. Which of the following statements is wrong?
(i) Single N–N bond is stronger than the single P–P bond.
(ii) PH_3 can act as a ligand in the formation of a coordination compound with transition elements.
(iii) NO_2 is paramagnetic.
(iv) Covalency of nitrogen in N_2O_5 is four
Solution:
Option (i) is the answer.

15. A brown ring is formed in the ring test for NO_3^- ion. It is due to the formation of
(i) $[Fe(H_2O)_5(NO)]^{2+}$
(ii) $FeSO_4.NO_2$
(iii) $[Fe(H_2O)_4(NO)]^{2+}$
(iv) $FeSO_4.HNO_3$
Solution:
Option (i) is the answer.

16. Elements of group-15 form compounds in +5 oxidation state. However, bismuth forms only one well-characterised compound in +5 oxidation state.

The compound is
(i) Bi_2O_5
(ii) BiF_5
(iii) $BiCl_5$
(iv) Bi_2S_5
Solution:
Option (ii) is the answer.

17. On heating ammonium dichromate and barium azide separately we get
(i) N_2 in both cases
(ii) N_2 with ammonium dichromate and NO with barium azide
(iii) N_2O with ammonium dichromate and N_2 with barium azide
(iv) N_2O with ammonium dichromate and NO_2 with barium azide
Solution:
Option (i) is the answer.

18. In the preparation of HNO_3, we get NO gas by catalytic oxidation of ammonia. The moles of NO produced by the oxidation of two moles of NH_3 will be _____.
(i) 2
(ii) 3
(iii) 4
(iv) 6
Solution:
Option (i) is the answer.

19. The oxidation state of a central atom in the anion of compound NaH_2PO_2 will be _ .
(i) +3
(ii) +5
(iii) +1
(iv) –3
Solution:
Option (iii) is the answer.

20. Which of the following is not tetrahedral in shape?
(i) NH_4^+
(ii) $SiCl_4$
(iii) SF_4
(iv) SO_4^{2-}
Solution:
Option (iii) is the answer.

21. Which of the following are peroxoacids of sulphur?
(i) H_2SO_5 and $H_2S_2O_8$
(ii) H_2SO_5 and $H_2S_2O_7$

(iii) $H_2S_2O_7$ and $H_2S_2O_8$
(iv) $H_2S_2O_6$ and $H_2S_2O_7$
Solution:
Option (i) is the answer.

22. Hot conc. H_2SO_4 acts as a moderately strong oxidising agent. It oxidises both metals and nonmetals. Which of the following element is oxidised by conc H_2SO_4 into two gaseous products?
(i) Cu
(ii) S
(iii) C
(iv) Zn
Solution:
Option (iii) is the answer.

23. A black compound of manganese reacts with a halogen acid to give greenish yellow gas. When an excess of this gas reacts with NH_3 an unstable trihalide is formed. In this process the oxidation state of nitrogen changes from _____.
(i) – 3 to +3
(ii) – 3 to 0
(iii) – 3 to +5
(iv) 0 to – 3
Solution:
Option (i) is the answer.

24. In the preparation of compounds of Xe, Bartlett had taken O_2+Pt F_6– as a base compound. This is because
(i) both O_2 and Xe has the same size.
(ii) both O_2 and Xe has the same electron gain enthalpy.
(iii) both O_2 and Xe has almost same ionisation enthalpy.
(iv) both Xe and O_2 are gases.
Solution:
Option (iii) is the answer.

25. In solid state PCl_5 is a _____.
(i) covalent solid
(ii) octahedral structure
(iii) ionic solid with $[PCl_6]^+$ octahedral and $[PCl_4]^-$ tetrahedra
(iv) ionic solid with $[PCl_4]^+$tetrahedral and $[PCl_6]^-$ octahedral
Solution:
Option (iv) is the answer.

26. Reduction potentials of some ions are given below. Arrange them in decreasing order of oxidising power.

Ion **Reduction potential E^o/V**	**ClO_4^-** **E^o = 1.19V**	**IO_4^-** **E^o = 1.65V**	**BrO_4^-** **E^o = 1.74V**

(i) ClO_4^- > IO_4^- > BrO_4^-
(ii) IO_4^- > BrO_4^- > ClO_4^-
(iii) BrO_4^- > IO_4^- > ClO_4^-
(iv) BrO_4^- > ClO_4^- > IO_4^-
Solution:
Option (iii) is the answer.

27. Which of the following is isoelectronic pair?
(i) ICl_2, ClO_2
(ii) BrO_2^-, BrF_2^+
(iii) ClO_2, BrF
(iv) CN^-, O_3
Solution:
Option (ii) is the answer.

II. Multiple Choice Questions (Type-II)
Note: In the following questions two or more options may be correct.

28. If chlorine gas is passed through a hot NaOH solution, two changes are observed in the oxidation number of chlorine during the reaction. These are _____ and _____.
(i) 0 to +5
(ii) 0 to +3
(iii) 0 to –1
(iv) 0 to +1
Solution:
Option (i) and (iii) are the answers.

29. Which of the following options are not in accordance with the property mentioned against them?

(i) F_2 > Cl_2 > Br_2 > I_2	**Oxidising power.**
(ii) MI > MBr > MCl > MF	**Ionic character of metal halide.**
(iii) F_2 > Cl_2 > Br_2 > I_2	**Bond dissociation enthalpy.**
(iv) HI < HBr < HCl < HF	**Hydrogen-halogen bond strength.**

Solution:
Option (ii) and (iii) are the answers

30. Which of the following is correct for P_4 the molecule of white phosphorus?
(i) It has 6 lone pairs of electrons.
(ii) It has six P–P single bonds.
(iii) It has three P–P single bonds.
(iv) It has four lone pairs of electrons.
Solution:
Option (ii) and (iv) are the answers.

31. Which of the following statements are correct?
(i) Among halogens, radius ratio between iodine and fluorine is maximum.
(ii) Leaving F—F bond, all halogens have weaker X—X bond than X—X' bond in interhalogens.
(iii) Among interhalogen compounds, the maximum number of atoms are present in iodine fluoride.
(iv) Interhalogen compounds are more reactive than halogen compounds.
Solution:
Option (i), (iii) and (iv) are the answers.

32. Which of the following statements are correct for SO_2 gas?
(i) It acts as a bleaching agent in moist conditions.
(ii) Its molecule has linear geometry.
(iii) It's dilute solution is used as a disinfectant.
(iv) It can be prepared by the reaction of dilute H_2SO_4 with metal sulphide.
Solution:
Option (i) and (iii) are the answers.

33. Which of the following statements are correct?
(i) All three N—O bond lengths in HNO_3 are equal.
(ii) All P—Cl bond lengths in PCl_5 molecule in the gaseous state are equal.
(iii) P_4 molecule in white phosphorus have angular strain therefore white phosphorus is very reactive.
(iv) PCl is ionic in the solid-state in which cation is tetrahedral and the anion is octahedral.
Solution:
Option (iii) and (iv) are the answers.

34. Which of the following orders are correct as per the properties mentioned against each?

(i) $As_2O_3 < SiO_2 < P_2O_3 < SO_2$	**Acid strength.**
(ii) $AsH_3 < PH_3 < NH_3$	**Enthalpy of vaporization.**
(iii) S < O < Cl < F	**More negative electron gain enthalpy.**
(iv) $H_2O > H_2S > H_2Se > H_2Te$	**Thermal stability.**

Solution:
Option (i) and (iv) are the answers.

35. Which of the following statements are correct?
(i) S–S bond is present in $H_2S_2O_6$
(ii) In peroxosulphuric acid (H_2SO_5) sulphur is in +6 oxidation state.
(iii) Iron powder along with Al_2O_3 and K_2O is used as a catalyst in the preparation of NH_3 by Haber's process.
(iv) Change in enthalpy is positive for the preparation of SO_3 by catalytic oxidation of SO_2
Solution:

Option (i) and (ii) are the answers.

36. In which of the following reactions conc. H_2SO_4 is used as an oxidising reagent?
(i) $CaF_2 + H_2SO_4 \rightarrow CaSO_4 + 2HF$
(ii) $2HI + H_2SO_4 \rightarrow I_2 + SO_2 + 2H_2O$
(iii) $Cu + 2H_2SO_4 \rightarrow CuSO_4 + SO_2 + 2H_2O$
(iv) $NaCl + H_2SO_4 \rightarrow NaHSO_4 + HCl$
Solution:
Option (ii) and (iii) are the answers.

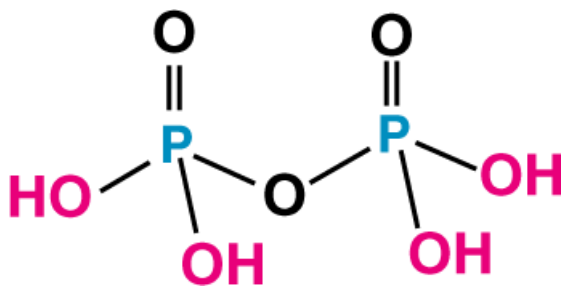
37. Which of the following statements are true?
(i) The only type of interactions between particles of noble gases is due to weak dispersion forces.
(ii) Ionisation enthalpy of molecular oxygen is very close to that of xenon.
(iii) Hydrolysis of XeF_6 is a redox reaction.
(iv) Xenon fluorides are not reactive.
Solution:
Option (i) and (ii) are the answers.

III. Short Answer Type

38. In the preparation of H_2SO_4 by Contact Process, why is SO_3 not absorbed directly in water to form H_2SO_4?
Solution:
Dissolution of SO_3 in water is highly exothermic. This leads to the formation of a mist of tiny droplets which are highly corrosive and they can even attack the lead pipelines (lead pipelines are used to cover the tower in contact process).

39. Write a balanced chemical equation for the reaction showing catalytic oxidation of NH_3 by atmospheric oxygen.
Solution:
$NH_3 + O_2 \rightarrow$(Pt gauze) $4NO + 6H_2O$
Pt gauze reacts as a catalyst, which is added to increase the rate of reaction.

40. Write the structure of pyrophosphoric acid.
Solution:



41. PH3 forms bubbles when passed slowly in water but NH3 dissolves. Explain why?
Solution:
PH3 does not form hydrogen bonding with water. Therefore it is not soluble in water and it escapes as a gas and forms bubbles whereas ammonia forms hydrogen bonding and soluble in water.

42. In PCl5, phosphorus is in sp3d hybridised state but all its five bonds are not equivalent. Justify your answer with reason.
Solution:
The size of axial bonds is greater than the size of equatorial bonds to overcome repulsion because the three equatorial bonds cause more repulsion. Therefore two axial P-Cl bonds are longer and different from equatorial bonds.

43. Why is nitric oxide paramagnetic in gaseous state but the solid obtained on cooling it is diamagnetic?
Solution:
Nitric acid in the gaseous state exists in monomer form. It consists of only one unpaired electron therefore it is paramagnetic. In solid-state, it exists as a dimer (N2O2). There is no unpaired electron in its dimer form, therefore it is diamagnetic.

44. Give the reason to explain why ClF3 exists but FCl3 does not exist.
Solution:
Chlorine has vacant d orbitals hence it can show an oxidation state of +3. Fluorine has no d orbitals, it cannot show a positive oxidation state. Fluorine shows only -1 oxidation state. Therefore FCl3 does not exists.

45. Out of H2O and H2S, which one has a higher bond angle and why?
Solution:
H2O has higher bond angle than H2S because as we move from oxygen to sulphur the size of the central atom increases and electronegativity decreases due to which bond pair goes away from the central atom which results in a decrease in bond pair repulsion and hence bond angle decreases.

46. SF6 is known but SCl6 is not. Why?
Solution:
Fluorine is the strongest oxidizing agent and it can oxidize sulphur to its maximum oxidation state +6 to

form SF_6. Chlorine is not a good oxidizing agent, it cannot oxidize sulphur to its maximum oxidation state. Chlorine can oxidize sulphur to only +4 oxidation state. Hence it can form SCl_4 but not SCl_6.

47. On reaction with Cl_2, phosphorus forms two types of halides 'A' and 'B'. Halide A is a yellowish-white powder but halide 'B' is a colourless oily liquid. Identify A and B and write the formulas of their hydrolysis products.
Solution:
Halide A is PCl_5 because it is a yellowish-white powder
Halide B is PCl_3 because it is a colourless oily liquid.
$PCl_3 + 3H_2O \rightarrow H_3PO_3 + 3HCl$
PCl_5 undergoes a violent hydrolysis
$PCl_5 + H_2O \rightarrow POCl_3 + 2HCl$
$PCl_5 + 4H_2O \rightarrow H_3PO_4 + 5HCl$

48. In the ring test of NO_3^- ion, Fe^{2+} ion reduces nitrate ion to nitric oxide, which combines with Fe^{2+} (aq) ion to form the brown complex. Write the reactions involved in the formation of a brown ring.
Solution:
$NO_3^- + 3Fe^{2+} + 4H^+ \rightarrow NO + 3Fe^{2+} + 2H_2O$
$[Fe(H_2O)_6]SO_4 + NO \rightarrow [Fe(H_2O)_5NO]SO_4$
Thus the complex formed is brown.

49. Explain why the stability of oxoacids of chlorine increases in the order given below:
$HClO < HClO_2 < HClO_3 < HClO_4$
Solution:
As the electronegativity of halogen decreases, the tendency of XO_3 group (X = halogens) to withdraw electrons of the O-H bond towards itself decreases and hence the acid strength of the perhalic acid decreases.

50. Explain why ozone is thermodynamically less stable than oxygen.
Solution:
Ozone is thermodynamically less stable because it decomposes into oxygen and this decomposition leads result in the liberation of heat, so its entropy is positive and free energy is negative.

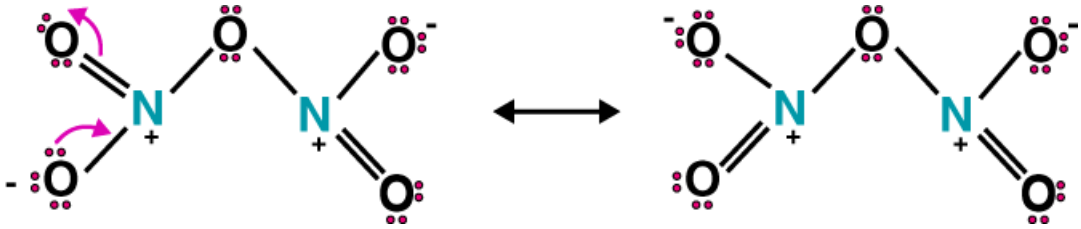
51. P_4O_6 reacts with water according to equation $P_4O_6 + 6H_2O \rightarrow 4H_3PO_3$.
Calculate the volume of 0.1 M NaOH solution required to neutralise the acid formed by dissolving 1.1 g of P_4O_6 in H_2O.
Solution:
$P_4O_6 + 6H_2O \rightarrow 4H_3PO_3$
$4H_3PO_3 + 8NaOH \rightarrow Na_2HPO_3 + 8H_2O$
Overall reaction :
$P_4O_6 + 8NaOH \rightarrow 4Na_2HPO_3 + 2H_2O$
Moles of $P_4O_6 = 1.1/220 = 0.5$
Acid formed by one mole of P_4o_6 requires = 8 mol
Acid formed by 0.005 mol of NaOH is present in 100mL solution
0.04 of NaOH is present in solution = 1000/0.1 *0.04 = 400mL

52. White phosphorus reacts with chlorine and the product hydrolyses in the presence of water. Calculate the mass of HCl obtained by the hydrolysis of the product formed by the reaction of 62 g of white phosphorus with chlorine in the presence of water.
Solution:
When white phosphorous reacts with chlorine:
$P_4 + 6Cl_2$ $4PCl_3$
$PCl_3 + 3H_2O$ $H_3PO_3 + 4HCl$]*4
$P_4 + 6Cl_2 + 12H_2O$ $4H_3PO_3 + 12HCl$
Moles of white P= 62/124 = 0.5mol
1mol of white P_4 produces HCl = 12mol
0.5 mol of white P_4 will produce HCl = 12*0.8 = 6mol
Mass of HCl = 6*36.5 = 219.0g

53. Name three oxoacids of nitrogen. Write the disproportionation reaction of that oxoacid of nitrogen in which nitrogen is in +3 oxidation state.
Solution:
There are three oxoacids of nitrogen.
Nitric acid(HNO_3), Nitrous acid (HNO_2), Hyponitrous acid ($H_2N_2O_1$)
+ 3 oxidation is shown by HNO_2, Therefore it undergoes disproportion reaction,

To calculate oxidation state, consider the oxidation state of N is x.
$3HNO_2$ $HNO_3 + 2NO + H_2O$
Oxidation state of nitrogen in HNO_2 is +3
X + (+1) +2*(-2) = 0, x= +3
Oxidation state of nitrogen in HNO_3 is +5
X + (+1) + 3*(-2) = 0, x= 5
Oxidation state of nitrogen in NO is +2
X + (-2) = 0, x =2
Therefore +3 oxidation state changes to +5 and +2 oxidation states.

54. Nitric acid forms an oxide of nitrogen on reaction with P_4O_{10}. Write the reaction involved. Also, write the resonating structures of the oxide of nitrogen formed.
Solution:
$P_4O_{10} + 4HNO_3$ $4HPO_3 + 2N_2O_5$
White phosphorous is very reactive as compared to red phosphorous due to angular strain in white phosphorous.



55. Phosphorus has three allotropic forms — (i) white phosphorus (ii) red phosphorus and (iii) black phosphorus. Write the difference between white and red phosphorus on the basis of their structure and reactivity.
Solution:
1. White phosphorus
(i) It is soft waxy solid having garlic order
(ii) It is poisonous
(iii) Has low melting point and boiling point because P4 molecules are held together by weak van der Waal forces of attraction

2. Red Phosphorus
(i) It is hard, crystalline, odourless solid
(ii) It is non-poisonous
(iii) It has a high melting point because of its polymeric structure.

56. Give an example to show the effect of concentration of nitric acid on the formation of oxidation product.
Solution:
$4Zn + 10HNO_3(dil) \rightarrow 4Zn(NO_3)_2 + NH_4NO_3 + 3H_2O$

$4Zn + 10HNO_3(dil) \rightarrow 4Zn(NO_3)_2 + N_2O + 5H_2O$

$4Zn + 4HNO_3(conc.) \rightarrow Zn(NO_3)_2 + NO_2 + 2H_2O$

57. PCl_5 reacts with finely divided silver on heating and a white silver salt is obtained, which dissolves on adding excess aqueous NH_3 solution. Write the reactions involved to explain what happens.
Solution:
PCl_5 reacts with silver to form white silver salt(AgCl). Which than dissolves on adding excess aqueous NH_3 to form a soluble complex
$PCl_5 + 2Ag \rightarrow 2AgCl$(white ppt.) $+ PCl_3$
$AgCl + 2NH_3(aq) \rightarrow [Ag(NH_3)]^+Cl^-$

58. Phosphorus forms a number of oxoacids. Out of these oxoacidsphosphinicacid has strong

reducing property. Write its structure and also write a reaction showing its reducing behaviour.
Solution:
$4AgNO_3 + H_3PO_2 + 2H_2O \rightarrow 4Ag + H_3PO_4 + 4HNO_3$
$2HgCl_2 + H_3PO_2 + 2H_2O \rightarrow 2Hg + H_3PO_4 + 4HCl$